

## AMENDMENTS TO THE SPECIFICATION

### In the Specification

Please replace paragraph [0036] beginning on page 15 with the following amended paragraph:

[0036] An optional analysis system 220, such as a spectrometer, is shown in the illustrated system, [[. The analysis system]] may exchange interrogation 221 and interrogation result signals 222 with the solution in the electroless electrolyte solution supply line. The analysis may measure a quantity that is expected to change over time in the solution, such as color, particle content, or spectroscopic absorption at a particular wavelength. This analysis may be used to assess and verify suitability of the solution and facilitate quality control.

Please replace paragraph [0037] beginning on page 15 with the following amended paragraph:

[0037] The solution in the electroless electrolyte solution supply line may be provided to a second fluid combination device, such as a reducing agent introduction tee 225. Then, often at a time proximate to the providing of this solution, a reducing agent or solution thereof 230 may be provided to the reducing agent introduction tee via a reducing agent supply line 231. The second tee may combine the solutions and often mix them. Then the tee may provide the resultant completed electroless deposition solution to the electroless deposition process 235, via an intervening electroless deposition solution supply line 226.

Please replace paragraph [0038] beginning on page 15 with the following amended paragraph:

[0038] **Figure 3** is a plot of spectroscopic absorption data for three electroless [[electrolyte]] electrolyte solutions on the y-axis versus age of the solutions in days on the x-axis, according to one embodiment of the invention. A first solution was prepared as

described in **Example 1** and contains a subset omitting tetramethylammonium hydroxide (TMAH). A second solution was prepared as described in **Example 2** and contains a subset omitting ammonium chloride (NH<sub>4</sub>Cl). A third solution was prepared similar to the first solution, except that instead of omitting TMAH, a sufficient amount of TMAH was included to adjust the pH to approximately 10. Accordingly the third solution contains the entire group.

Please replace paragraph [0039] beginning on page 16 with the following amended paragraph:

[0039] Spectroscopic absorptions for the first, second, and third solutions are indicated by the lines marked respectively with open triangles ( $\Delta$ ), open circles (O), and ~~[[solid squares (■)]]~~ open squares ( $\square$ ). The absorption data reflect maximum absorptions for the three solutions measured with an ultraviolet-visible spectrometer at a wavelength in the range of approximately 510-530 nanometers (nm). At day zero the absorption for the first solution was approximately 0.4 and for the second and third solutions was approximately 1. The ~~[[differences]]~~ difference is due to difference in compositions.

Please replace paragraph [0041] beginning on page 16 with the following amended paragraph:

[0041] In contrast, the line representing the ~~[[third solution (■)]]~~ third solution ( $\square$ ) steadily increased in y-value from an initial value of approximately 1 to a final value on day twenty of approximately 2.5. This significant increase in the absorption -- about 250% -- is indicative of significant transformation of the solution due to chemical reactions. Over the twenty-day period the color of the third solution changed and minute particles formed. No such color change or particle formation was observed in the first and second solutions. These experimental results are believed to indicate relative chemical stability in the first and second solutions, according to an embodiment of the invention, compared to the third solution.